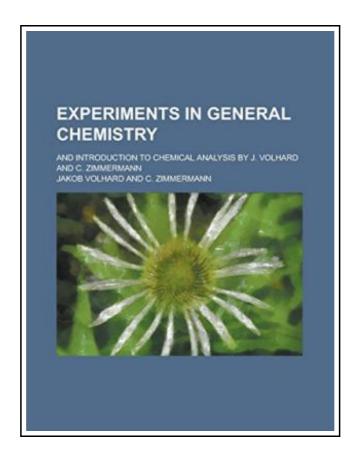
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Rarebooksclub.com, United States, 2013. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****. This historic book may have numerous typos and missing text. Purchasers can usually download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1887 edition. Excerpt: .in NHs and NH4Cl, easily soluble in mineral acids and in acetic acid.J If the Jr moist sulphide is exposed to the air it becomes brown, and on treating with water a part dissolves as manganous salt (P). 555 /llangarzic oxide (anhydrous as braunite, MmOa; with water of crystallization as manganite, MmOa---HtO) is a very weak base whose unstable salts are cherry red or purple in solution. The phosphate is the most stable; prepare it as follows: add a few drops of NOBH to a manganous salt, add conc. phosphoric acid, and boil in a porcelain dish till the nitric acid begins to evaporate (reaction for Mn); MntOa colors the microcosmic bead amethyst red. All Mn compounds form this bead in the oxidizing flame; in the reducing flame the bead is colorless (P).5194 Hydroxides of manganese dioxide, MnOe, are formed when manganous salts are precipitated by alkali hydroxide in presence of an oxidizing agent; sodium hypochlorite (?)j, jBr or Cl (?)j, epotassium permanganate, form thus with manganous salt solutions brown precipitates of perhydroxides, Mn(OH)t or MnOtHt. From an acidulated (not strongly acid) solution of a manganous salt to which sodium acetate has been added, on warming with Br water (?) i1 Mn is precipi and all other oxides of Mn riclz in O (?)J,6behave with conc. SO4H2, with HCI, with dilute acids and reducing agents, like the peroxide; all are reduced to manganous salt, giving up their oxygen as oxygen, or yielding an equivalent...

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